**CHEMISTRY 2022-23 January 4, 2023**

**Today’s Agenda (Day 79)**

1. HOUSEKEEPING ITEMS

**🡪**  BRING:

1. Homework Check:

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1. Class Activity:

🡪DAY 2: Chapter 11 PPT Review

1. ~~Section 11.1 – Defining Stoichiometry~~
2. **Section 11.2 – Stoichiometric Calculations**
3. Section 11.3 – Limiting Reactants
4. Section 11.4 – Percent Yield

HOMEWORK:

* READ: Chapter 11 – Stoichiometry
* COMPLETE:
* STUDY: Chapter 11 Test

REMINDERS:

* QUIZ:Balancing Equations & Naming Compounds – Jan. 10
* TEST: **Ch 11 🡪 Jan. 12**

**CHEMISTRY 2022-23 LAUNCH LAB**

**CHAPTER 11 LAUNCH LAB – What Evidence Can You Observe That a Reaction is Taking Place?**

During a chemical reaction, reactants are consumed as new products are formed. Often, there are several telltale signs that a chemical reaction is taking place.

**Procedure **

1. Read and complete the lab safety form.

2. Use a **10-mL graduated cylinder** to measure out 5.0 mL **0.01M potassium permanganate** (KMnO4). Add the solution to a **100-mL beaker**.

3. Clean and dry the graduated cylinder, and then use it to measure 5.0 mL **0.01M sodium hydrogen sulfite** solution (NaHSO3). Slowly add this solution to the beaker while stirring with a stirring rod. Record your observations.

4. Repeat Step 3 until the (KMnO4) solution in the beaker turns colorless. Stop adding the NaHSO3 solution as soon as you obtain a colorless solution. Record your observations.

**Analysis**

1. Identify the evidence you observed that a chemical reaction was occurring.

2. Explain why slowly adding the NaHSO3 solution while stirring is a better experimental technique than adding 5.0 mL of the solution all at once.

**Inquiry**

Would anything more have happened if you continued to add NaHSO3 solution to the beaker? Explain.

**CHEMISTRY 2022-23 PRACTICE PROBLEM**

**CHAPTER 11 – Stoichiometry**

**Practice Problems 11.1 –** Interpreting Chemical Equations

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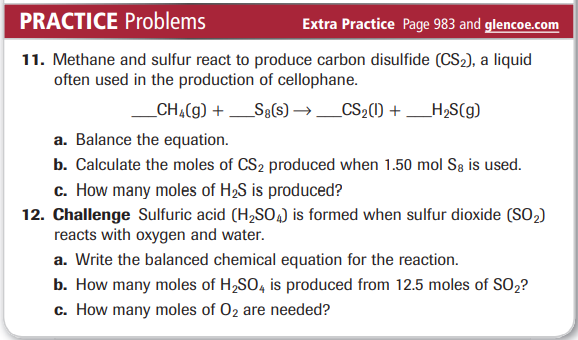
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**Practice Problems 11.1 –** Mole Ratios

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**Practice Problems 11.2 –** Mole-to-Mole Stoichiometry



**Practice Problems 11.3 –** Mole-to-Mass Stoichiometry

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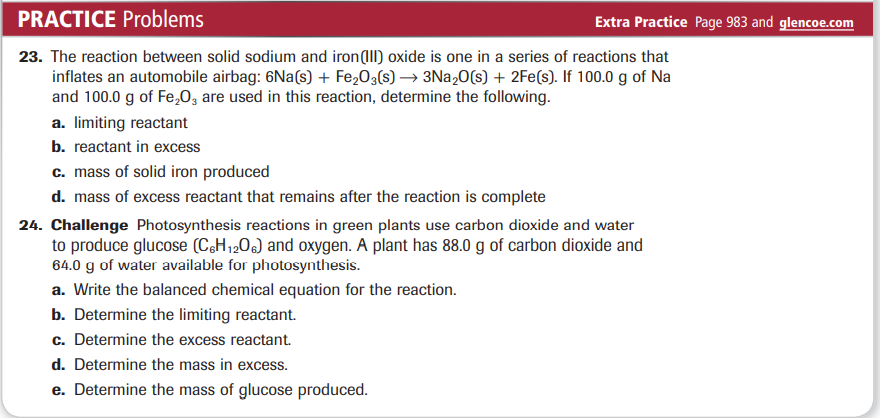
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**Practice Problems 11.4 –** Mass-to-Mass Stoichiometry

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**Practice Problems 11.5 –** Determining the Limiting Reactant



**Practice Problems 11.6 –** Percent Yield

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**CHEMISTRY 2022-23 MINI LAB**

**CHAPTER 11 MINI LAB – Apply Stoichiometry**

**How much sodium carbonate (Na2CO3) is produced when baking soda decomposes?**

Baking soda is used in many baking recipes because it makes batter rise, which results in a light and fluffy texture. This occurs because baking soda, sodium hydrogen carbonate (Na2CO3), decomposes upon heating to form carbon dioxide gas according to the following equation.

2 Na2CO3 → Na2CO3 + CO2 + H2O

**Procedure **

1. Read and complete the lab safety form.

2. Create a data table to record your experimental data and observation.

3. Use a **balance** to measure the mass of a clean, dry **crucible**. Add about 3.0 g of **sodium hydrogen carbonate** (Na2CO3) and measure the combined mass of the crucible and Na2CO3. Record both masses in your data table and calculate the mass of the Na2CO3.

4. Use this starting mass of Na2CO3 and the balanced chemical equation to calculate the mass of Na2CO3 that will be produced.

5. Set up a **ring stand** with a **ring** and **clay triangle** for heating the crucible.

6. Heat the crucible with a **Bunsen burner**, slowly at first and then with a stronger flame, for 7–8 min. Record your observations during the heating.

7. Turn off the burner and use **crucible tongs** to remove the hot crucible.

**WARNING: Do not touch the hot crucible with your hands.**

8. Allow the crucible to cool, and then measure the mass of the crucible and Na2CO3.

**Analysis**

1. Describe what you observed during the heating of the baking soda.

2. Compare your calculated mass of Na2CO3 with the actual mass you obtained from the experiment.

3. Calculate Assume that the mass of Na2CO3 that you calculated in Step 4 is the accepted value for the mass of product that will form. Calculate the error and percent error associated with the experimentally measured mass.

4. Identify sources of error in the procedure that led to errors calculated in Question 3.