**CHEMISTRY 2022-23 September 26, 2022**

**Today’s Agenda (Day 28)**

1. HOUSEKEEPING ITEMS

**🡪**

1. Homework Check:

🡪 Chapter 5 Practice Problems

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1. Class Activity:

🡪**CHEMLAB 4: Model Atomic Mass** – see p. 2 of document; work in PAIRS or THREES

🡪 TUESDAY: Chapter 5 PPT Review – DAY 2

1. **Section 5.1 – Light and Quantized Energy**
2. Section 5.2 – Quantum Theory and the Atom
3. Section 5.3 – Electron Configuration

HOMEWORK:

* READ: Chapter 5 – Electrons in Atoms
* COMPLETE: Chapter 6 Vocabulary
* STUDY: Chapter 5 Test, Chapter 5 & 6 Vocabulary

CHAPTER 5

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| Amplitude | atomic emission spectrum | atomic orbital | Aufbau principle | de Broglie equation | electromagnetic radiation |
| electromagnetic spectrum | electron configuration | electron-dot structure | energy sublevel | Frequency | ground state |
| Heisenberg uncertainty principle | Hund's rule | Pauli exclusion principle | photoelectric effect | Photon | Planck's constant |
| principal energy level | principal quantum number | Quantum | quantum mechanical model of the atom | quantum number | valence electron |
| wavelength |  |  |  |  |  |

CHAPTER 6

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| actinide series | alkali metal | alkaline earth metal | Electronegativity | Group | Halogen |
| inner transition metal | Ion | ionization energy | lanthanide series | Metal | Metalloid |
| noble gas | Nonmetal | octet rule | Period | periodic law | representative element |
| transition element | transition metal |  |  |  |  |

REMINDERS:

* ~~Chapter 5 Practice Problems – Sept. 27~~
* TEST: **Ch 5** 🡪 **Sept. 29**
* Chapter 6 Vocabulary – Oct. 11
* QUIZ: **Ch 5 & 6 Vocabulary 🡪 Oct. 13**
* TEST: **Ch 6 🡪 Oct. 18**

**CHEMISTRY 2022-23 CHEMLAB**

**CHEMLAB 4: Model Atomic Mass** A picture containing indoor, sliced

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**Background**: Most elements in nature occur as a mixture of isotopes. The weighted average atomic mass of an element can be determined from the atomic mass and the relative abundance of each isotope. In this activity, you will model the isotopes of the imaginary element “Snackium.” The measurements you make will be used to calculate a weighted average mass that represents the average atomic mass of “Snackium.”.

**Question**: How are the atomic masses of the natural isotopic mixtures calculated?

**Materials**

balance

calculator

bag of snack mix

**Safety Precautions** 

WARNING: Do not eat food used in lab work.

**Procedure**

1. Read and complete the lab safety form.

2. Create a table to record your data. The table will contain the mass and the abundance of each type of snack present in the mixture.

3. Open your snack-mix bag. Handle the pieces with care.

4. Organize the snack pieces into groups based on their types.

5. Count the number of snack pieces in each of your groups.

6. Record the number of snack pieces in each group and the total number of snack pieces in your data table.

7. Measure the mass of one piece from each group and record the mass in your data table.

8. Cleanup and Disposal Dispose of the snack pieces as directed by your teacher. Return all lab equipment to its designated location.

**Analyze and Conclude**

1. Calculate Find the percent abundance of the pieces by dividing the individual-piece quantity by the total number of snack pieces.

2. Calculate Use the isotopic percent abundance of the snack pieces and the mass to calculate the weighted average atomic mass for your element “Snackium.”

3. Interpret Explain why the weighted average atomic mass of the element “Snackium” is not equal to the mass of any of the pieces.

4. Peer Review Gather the average atomic mass data from other lab groups. Explain any differences between your data and the data obtained by other groups.

5. Apply Why are the atomic masses on the periodic table not expressed as whole numbers like the mass number of an element?

6. Research Look in a chemical reference book to determine whether all elements in the periodic table have isotopes. What is the range of the number of isotopes chemical elements have?

7. Error Analysis What sources of error could have led the lab groups to different final values? What modifications could you make in this investigation to reduce the incidence of error?

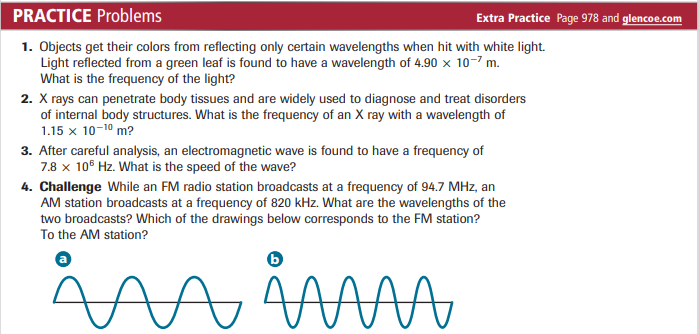
**INQUIRY EXTENSION**

Predict Based on your experience in this lab, look up the atomic masses of several elements on the periodic table and predict the most abundant isotope for each element.

**CHEMISTRY 2022-23 PRACTICE PROBLEMS**

**CHAPTER 5 – Structure of the Atom**

**Practice Problems 5.1 –** Calculating Wavelength of an EM Wave



**Practice Problems 5.2 –** Calculate the Energy of a Photon

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**Problem-Solving Lab –** Interpret Scientific Illustrations

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**Practice Problems 5.3 –** Filling Atomic Orbitals

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**Practice Problems 5.3 –** Electron-Dot Structures

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